Sınaq	7					
	Kimya					
61	C	**71**	D	**81**	B
62	E	**72**	E	**82**	C
63	A	**73**	B	**83**	245
64	D	**74**	D	**84**	3
65	E	**75**	C	**85**	24
66	D	**76**	B	**86**	23
67	B	**77**	D	**87**	1B2C3A
68	A	**78**	C		
69	D	**79**	E		
70	A	**80**	E		

88

$\gamma = 2$

1. $a = \upsilon_{t_2}$

$$\upsilon_{t_2} = \upsilon_{t_1} \cdot \gamma^{\frac{t_2 - t_1}{10}} = 0{,}02 \cdot 2^{\frac{85-65}{10}} \; 0{,}02 \cdot 2^2 = 0{,}08 \, \text{mol/(l} \cdot \text{san)}$$

$b = \tau_2$

2. $$\frac{\upsilon_{t_2}}{\upsilon_{t_1}} = \frac{\tau_1}{\tau_2} = \frac{0{,}08}{0{,}02} = \frac{120}{b} \qquad b = 30 \, \text{san}$$

89

2x q **32y q**

xH_2 **yO_2**

x + y = 20 ikiyə vursaq 2x + 2y = 40 30y = 120

23x + 32y = 160 2x + 32y = 160 **y** = 4 mol, onda, **x** =16 mol

$2H_2 + O_2 \longrightarrow H_2O$

baş: 16 4

sərf: 8 4

artıq: 8 0

Artıq qalan qaz: **H_2** və onun mol miqdarı: **8 mol**.

90

1. 0,1 mol **(Y_3N_2)** -------- 10 q

1 mol **(Y_3N_2)** -------- ***M*** ***M*(Y_3N_2)** = 100 q

Deməli ***M*$_r$(Y_3N_2) = *M*$_r$(XCO_3)** bərabərdirsə, onda ***M* (Y_3N_2) = *M* (XCO_3) = 100 q** olacaq.

***M*(XCO_3) = 100 q**

X + 12 + 48 = 100

X = 40

$$\omega_X = \frac{m(X)}{M(XCO_3)} \cdot 100\% = \frac{40}{100} \cdot 100\% = 40\%$$

2.

***M*(Y_3N_2)** = 100 q

3Y + 28 = 100

3Y = 72

Y = 24

$$\omega_Y = \frac{m(Y)}{M(Y_3N_2)} \cdot 100\% = \frac{72}{100} \cdot 100\% = 72\%$$