

Sınaq		8			
		Kimya			
61	E	71	D	81	E
62	B	72	C	82	C
63	A	73	A	83	34
64	C	74	B	84	23
65	E	75	E	85	960
66	C	76	B	86	24
67	B	77	E	87	1B2CD3AE
68	D	78	A		
69	C	79	D		
70	A	80	D		

88	$\omega = \frac{m_{\text{mad}}}{m_{\text{meh}}} \cdot 100\% = \frac{40}{160 + 40} \cdot 100\% = 20\%$ $K_h = \frac{m_{\text{mad}}}{m_{\text{su}}} \cdot 1000 = \frac{40}{160} \cdot 1000 = 250 \text{ q/l}$
89	$98 + 81 = M_r + 18$ $\text{Cu(OH)}_2 + \text{HBr} \rightarrow \text{Cu(OH)Br} + \text{H}_2\text{O}$ $98 + 81 = M_r + 18$ $X \rightarrow M_r = 161$ <p>Y → mis (II) hidrokso bromid (Cu(OH)Br)</p>
90	<p>1. $2\text{Ca} + \text{O}_2 \rightarrow 2\text{CaO}$</p> <p>2. $2 \cdot 40 \quad 32\text{q}$ $2\text{Ca} + \text{O}_2 \rightarrow 2\text{CaO}$ $\frac{40}{80} = 0,5 \quad \frac{40}{32} > 0,5$ Nisbəti kiçik olan tam sərf olunur. (hesablama Ca görə aparılır)</p> $2 \cdot 40 \text{ q} \text{ ----- } 2 \cdot 56 \text{ q}$ $2\text{Ca} + \text{O}_2 \rightarrow 2\text{CaO}$ $40 \text{ q} \text{ ----- } a$ $a = \frac{40 \cdot 2 \cdot 56}{2 \cdot 40} = 56 \text{ q}$